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Chapter 7 The
chemical formula for
an ionic compound
represents one formula
unit - the simplest ratio
of the compound's
positive ions (cations)
and its negative ions
(anions). Monatomic
cations are identified
simply by the
element's name.

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formula. subscript of
each element reveals
number of atoms
contained in a single
molecule of a certain
compound.

monoatomic ions. ions
formed from a single
atom; remember to
change nonmetals to
-ide ending. oxyanions.

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Vocabulary. STUDY.
PLAY. Formula mass.
The formula unit of a
molecule or ion is the
sum of the average
atomic masses of all
atoms represented in

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CHAPTER 7 REVIEW

Chemical Formulas and
Chemical Compounds

SECTION 1 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

c In a Stock system
name such as iron (III)
sulfate, the Roman
numeral tells us (a)
how many atoms of Fe
are in one formula unit.

(b) how many sulfate

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ions can be attached to
the iron atom.

Answers
**Holt Modern
Chemistry Chapter 7
Section 1 Review
Answers**

7H 5(NO 2) 3. 4
elements a. How many
elements make up this
compound? 6 oxygen
atoms b. How many
oxygen atoms are
present in one
molecule of C 7H 5(NO
2) 3? 21 atoms c. How
many atoms in total

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are present in one molecule of $C_7H_5(NO_2)_3$?

4.2 10^{24} atoms

d. How many atoms are present in a sample of 2.0×10^{23} molecules of $C_7H_5(NO_2)_3$?

7 Chemical Formulas and Chemical Compounds

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• CHAPTER 7

HOMEWORK 7-3

(pp.211-213)

VOCABULARY Write the prefix that corresponds to each number. Try

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not to use the book.

Prefix Number Number

Prefix 1 6 2 7 3 8 4 9 5

10 GRAPHIC

ORGANIZER Complete the chart by supplying the missing information.

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CHAPTER 7

CHAPTER 7 REVIEW

Chemical Formulas and
Chemical Compounds

MIXED REVIEW SHORT

ANSWER Answer the following questions in

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the space provided. 1.
Write formulas for the
following compounds:
CuCO₃ a. copper(II)
carbonate Na₂SO₃ b.
sodium sulfite (NH₄)
₃PO₄ c. ammonium
phosphate SnS₂ d.
tin(IV) sulfide HNO₂ e.
nitrous acid 2.

7 Chemical Formulas and Chemical Compounds

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