

Chapter 18 Chemical Equilibrium Answers Section 3

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Chapter 18 Chemical Equilibrium Answers

$\text{H}_2(\text{g}) + \text{Br}_2(\text{g}) \leftrightarrow 2\text{HBr}(\text{g})$ $K = 4.0 \times 10^{-2}$ $K = 3.5 \times 10^4$ Chapter 18 Chemical Equilibrium 18.2 Shifting Equilibrium The Effect of Change in Pressure: Ways to Change Pressure Add or remove a gaseous reactant or product Add an inert gas (one not involved in the reaction) An inert gas increases the total pressure but has no effect on the concentrations or partial pressures of the reactants or products The Effect of Change in Pressure Change the volume of the container When the volume of the ...

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different products represented.

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Reversible reaction. a chemical reaction in which products can react to re-form the reactants. Chemical equilibrium. when the rate of the forward reaction equals the rate of the reverse reaction and the concentrations of the products and the reactants remain unchanged.

Chapter 18: Chemical Equilibrium vocabulary Flashcards

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Chapter 18 Chemical Equilibrium Mixed Review Answers

Modern Chemistry 143 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. Write the equilibrium expression for the following hypothetical equation: $3A(aq) + B(aq) \rightarrow 2C(aq) + 3D(aq)$ ← 2. a. Write the appropriate chemical equilibrium expression for each of the

CHAPTER 18 REVIEW Chemical Equilibrium

Chemistry: Chapter 18- Chemical Equilibrium. Equilibrium. Value of K_c . $K_c < 1$. $K_c > 1$. when the rates of opposing actions are the same. represents the reaction progression... -all K_c 's are dependent on... reactants of the reaction are favored. products of the reaction are favored.

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sometimes referred to as the chemical-equilibrium expression. The H₂, I₂, HI Equilibrium System Consider the reaction between H₂ and I₂ vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing

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the rate at which the violet color of the iodine vapor diminishes, as $[C]^x[D]^y [A]^n[B]^m$ 556 CHAPTER 18

CHAPTER 18 Chemical Equilibrium

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Chapter 18: Chemical Equilibrium

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Chemical Equilibrium Important Questions And Answers

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